

WS 4.6 Types of Reactions

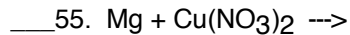
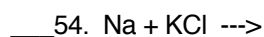
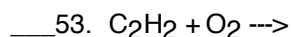
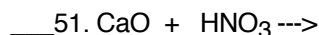
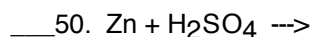
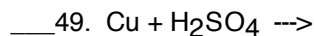
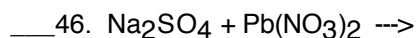
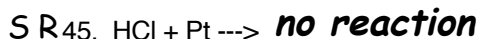
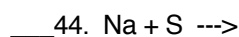
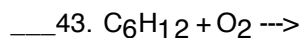
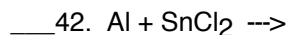
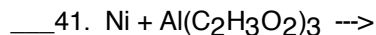
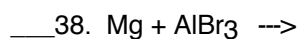
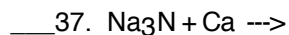
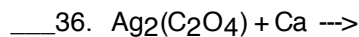
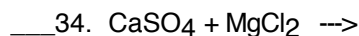
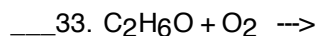
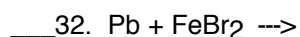
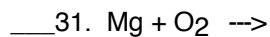
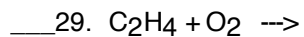
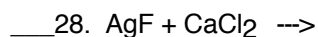
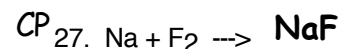
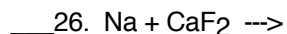
Complete the reactions by writing the products. *Remember* : when you form an element, don't forget about the 7 diatomic gases (N₂, O₂, F₂, H₂, Cl₂, Br₂, I₂) & to balance ionic formulas (drop & swap)

Composition Reactions	Decomposition Reactions	Combustion Reactions
1. Na + Cl ₂ --->	6. MgO --->	11. CH ₄ + O ₂ --->
2. K + O ₂ --->	7. AlCl ₃ --->	12. C ₅ H ₁₂ + O ₂ --->
3. H ₂ + F ₂ --->	8. H ₂ O --->	13. O ₂ + C ₆ H ₆ --->
4. Li + N ₂ --->	9. Na ₂ CO ₃ --->	14. C ₂ H ₅ OH + O ₂ --->
5. Mg + O ₂ --->	10. NF ₃ --->	15. C ₁₂ H ₂₂ O ₁₁ + O ₂ --->
Double Replacement Reactions	Single Replacement Reactions	
16. CaCl ₂ + Al ₂ O ₃ --->	21. AgCl + Mg --->	
17. LiCl + Pb(NO ₃) ₂ --->	22. Ca + FeF ₃ --->	
18. Na ₂ SO ₄ + CaCl ₂ --->	23. HCl + Al --->	
19. HCl + K ₃ PO ₄ --->	24. KBr + F ₂ --->	
20. HBr + Ca(OH) ₂ --->	25. Cl ₂ + AlI ₃ --->	

Determine the products & identify the type of reaction using these abbreviations in the spaces at left:

(**CP**=composition, **DC**=decomposition, **SR**=single replacement, **DR**=double replacement, **CB**=combustion)

*** Don't forget to check the activity series for single replacement reactions. There will be 5 more NR's ***



Li
K
Ca
Na
Mg
Al
Mn
Zn
Cr
Fe
Cd
Co
Ni
Sn
Pb
H
Cu
Hg
Ag
Pt
Au